

# Novel Graphene Material for High Energy Storage Supercapacitors

11-14<sup>th</sup> October 2022  
ESA/ESTEC, Noordwijk, The Netherlands

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## ABSTRACT

Supercapacitors have attracted great interest due to their high energy density, practically unlimited charging/discharging number of cycles and sustainability. However, compared to battery technologies their energy densities still remain lower. In the last decade, supercapacitors with an energy content of  $\sim 110$  Wh/L at a power of  $\sim 1$  kW/L were developed by leveraging the open framework structure of graphene-related architectures. The research described in this paper report a novel method of high energy density graphene materials prepared based on fluorographene.

The best performing highly nitrogen doped ( $\sim 16\%$ ) graphene (GN3), with diamond-like bonds and an ultra-high mass density of  $2.8$  g cm<sup>-3</sup>, is an excellent host for the ions, delivering unprecedented energy densities of up to  $200$  WhL<sup>-1</sup> at a power of  $2.6$  kW L<sup>-1</sup> and  $143$  Wh L<sup>-1</sup> at  $52$  kW L<sup>-1</sup>. These findings open a route to materials whose properties may enable a transformative improvement in the performance of supercapacitor components. The paper explains the technology background and features of this high-energy graphene-based supercapacitors, their challenges, and potentials for future applications.

## INTRODUCTION

The pursuit of materials for efficient electricity storage is one of the key challenges for today's science. The ever-increasing mobility and number of electronic devices, along with ongoing efforts to reduce fossil fuel consumption, will continue to spur global demand for affordable, reliable and sustainable energy. The goal is not only to continuously increase battery capacity but also to look for other efficient non-lithium-based ways of storing electricity. Carbon-based supercapacitors are beginning to be an attractive alternative for energy storage, mainly due to their safety, long-life and extraordinary ability to sustain millions of charging cycles without losing capacity.

Current research in SC is primarily focused on highly porous, high surface area activated carbons (i.e.,  $2000$  m<sup>2</sup>g<sup>-1</sup>), since the energy storage mechanism is based on the formation of the electrostatic double layer at the interface between the electrode material and the electrolyte. Such carbons have a very low density (c.a.  $0.3$ - $0.5$  gcm<sup>-3</sup>), leading to a large empty space in the electrode which is then flooded with electrolyte. Thus, the electrode not only consumes large volume but also gets heavier, while the electrolyte-filled space does not add to capacitance. Therefore, the highest so far obtained volumetric energy densities (parameter of critical importance for practical applications) of carbon-based SCs have been obtained from densely packed graphene architectures, ranging between  $80$  to  $100$  Wh L<sup>-1</sup>. These values are  $\sim 5$  times higher than commercial high performance active carbons and they can thus compete with lead acid and metal hydride batteries. Theoretically, though, graphene could potentially reach  $300$  WhL<sup>-1</sup>.

## Technology Background

Supercapacitors, also known as EDLCs (electric double-layer capacitor) or ultracapacitors, differ from regular capacitors in that they can store tremendous amounts of energy. There are three types of supercapacitors, based on the design of the electrodes (Figure 2.) and dominant storage mechanism (Figure 1).

In the case of batteries, a liquid electrolyte changes the chemical states of ions in order to operate, while a supercapacitor stores the ions on the surface of its electrodes in the form of static electricity. This is, in short, the main difference between the two. It means that a capacitor is able to deliver energy very quickly in big bursts and recharge almost as rapidly.

Pseudocapacitors are EDLCs that mainly store electrical energy Faradaically by electron charge transfer between electrode and electrolyte. There is no chemical reaction, battery Faradaic action. This is accomplished through electrosorption, reduction-oxidation reactions (redox reactions), and intercalation (electrode penetration) processes that are not in the battery electrode that is the basis of battery-supercapacitor hybrids. All three are termed pseudocapacitance. A pseudocapacitor is an inseparable part of an EDLC (supercapacitor). It forms together with an electric double-layer capacitor (EDLC) to create a supercapacitor. When an EDLC is designed with materials conferring a lot of pseudocapacitance, rather than minimizing to give more ideal capacitor behavior it is called a pseudocapacitor.

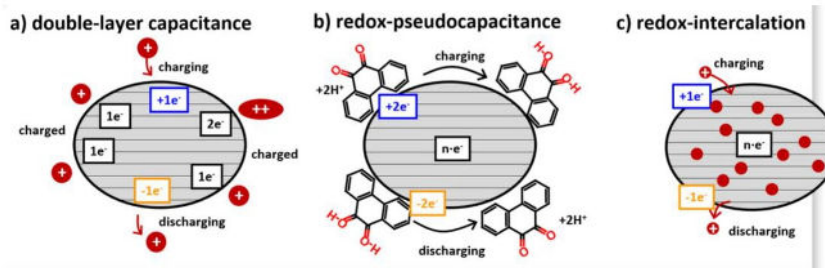


Figure 1. Charge / energy storage type mechanisms; source: Leibnitz Institute of New Materials

A basic capacitor usually consists of two metal plates, separated by an insulator (like air or a plastic film). During charging, electrons accumulate on one conductor and depart from the other. One side gains a negative charge while the other side builds a positive one. The insulator disturbs the natural pull of the negative charge towards the positive one, and that tension creates an electric field. Once electrons are given a path to the other side, discharge occurs. Supercapacitors contain two metal plates, coated with a porous material known as activated carbon. They are immersed in an electrolyte made of positive and negative ions dissolved in a solvent. One plate is positive, and the other is negative.

During charging, ions from the electrolyte accumulate on the surface of each carbon-coated plate. Supercapacitors also store energy in an electric field that is formed between two oppositely charged particles, only they have the electrolyte in which an equal number of positive and negative ions is uniformly dispersed. Thus, during charging, each electrode ends up having two layers of charge coating (electric double-layer).

Supercapacitors can be viewed as a kind of cross between the electrolyte-based battery and the capacitor. Like a capacitor, the ions of a supercapacitor are stored on the surface of the electrodes in the form of static electricity. However, it differs from traditional capacitors in that an electrolyte is used to provide the ions to the electrodes during charging.

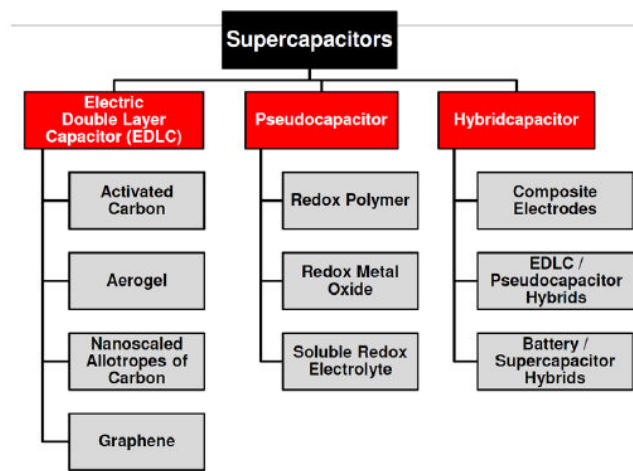


Figure 2. Types of SCs based on the design of electrodes.

Table 1. below summarizes the key characteristics of SCs in relation to capacitors and batteries.

Parameter	Aluminium electrolytic capacitors	Supercapacitors			Lithium-Ion batteries
		Double-layer capacitors for memory backup	Supercapacitors for power applications	Pseudo and Hybrid capacitors (Li-Ion capacitors)	
Temperature range (°C)	-40 to +125	-20 to +70	-20 to +70	-20 to +70	-20 to +60
Cell voltage (V)	4 to 550	1.2 to 3.3	2.2 to 3.3	2.2 to 3.8	2.5 to 4.2
Charge/discharge cycles	unlimited	10 <sup>5</sup> to 10 <sup>6</sup>	10 <sup>5</sup> to 10 <sup>6</sup>	2x10 <sup>4</sup> to 10 <sup>5</sup>	500 to 10 <sup>4</sup>
Capacitance range (F)	≤1	0.1 to 470	100 to 12000	300 to 3300	-
Energy density (Wh/kg)	0.01 to 0.3	1.5 to 3.9	4 to 9	10 to 15	100 to 265
Power density (kW/kg)	>100	2 to 10	3 to 10	3 to 14	0.3 to 1.5
Self discharge time at room temperature	Short (days)	Middle (weeks)	Middle (weeks)	Long (month)	Long (month)
Efficiency (%)	99	95	95	90	90
Life time at room temperature (years)	>20	5 to 10	5 to 10	5 to 10	3 to 5
	Capacitor	←Capacitor-like		Battery-like→	Battery

Table 1. Key characteristics of capacitors, supercapacitors and batteries

Supercapacitors are currently mostly used in conjunction with batteries, to give a quick burst of energy in applications that require both continuous energy and momentary energy bursts. If research activity will pay off and researchers find a way to increase the energy density of supercapacitors, they could pose an appealing alternative to traditional batteries for powering electric vehicles (EVs) or enable new applications in conjunction with the next generation of batteries.

SCs belong to the group of energy storage devices that display significant advantages over batteries or fuel cells, such as high-power density, extralong cycle life, and small size. SCs can store more energy than conventional capacitors, however, the energy which they can store is approximately one order of magnitude lower than that of batteries.

For instance, commercially available porous activated carbons (AC) have a typical value of energy density close to 10 Wh kg<sup>-1</sup> (15 Wh L<sup>-1</sup>), while lead-acid batteries have 20 - 35 Wh kg<sup>-1</sup> (40-80 Wh L<sup>-1</sup>) and Li-ion batteries reach up to ~200 Wh kg<sup>-1</sup> (~350 Wh L<sup>-1</sup>). However, Li-ion batteries suffer from long charging time (far non-competitive to conventional refuelling), and irreversible changes, which significantly limit their life in charging-discharging (c-d) cycles. In practice, this leads to a significant drop in energy density after a few hundreds of c-d cycles<sup>1</sup>.

## DEVELOPMENT OF HIGH-ENERGY GRAPHENE ELECTRODE SUPERCAPACITORS

EDLCs currently mostly use as electrodes Activated Carbon or Carbon Nanotubes (CND). The potential benefits of graphene electrodes vs carbon includes:

- More cost-efficient, enormous surface area (2630 m<sup>2</sup> g<sup>-1</sup>) and higher power density.
- 10x more conductive.
- Theoretical max capacitance of 2D graphene is 550 F g<sup>-1</sup> (~200 F g<sup>-1</sup> achieved in practice).
- Easy team up with various other nanomaterials, prominently carbon nanotubes (CNTs), to create low cost, lightweight and high-performance SCs.

We have considered several design strategies how to boost graphene-based electrode material charge storage potential such as:

- Increase the surface area (microporous 3D structure).
- Reduce restacking.
- Increase the packing density and conductivity.
- Accomplish defect control.
- Functionalization and hybridization of electrode materials.

Based on the above strategies we set the aim of the research into two directions how to boost the capacitance and energy density of graphene-based electrode materials:

**#1** Mounting of longer functional groups perpendicular to the graphene surface to further enhance capacitance.

**#2** Doping with heteroatoms such as N-doping, that increases the electronic conductivity and improves the solid-electrolyte interface, allowing solvated ions to better infiltrate the pores of the electrode – increasing both electrostatic and pseudo-capacitance mechanism.

The proposed way to meet the objectives was to use covalent functionalization of graphene and N-doping, with fluorographene proposed as the precursor for functionalization. This process has been proved in previous work<sup>2,4</sup> as a low cost, ideal to create high-density matching material with graphene-like monolayer.

The graphite fluoride material is an industrial lubricant available on the market in tonnes, which increases its potential commercial availability and smooth transition into mass manufacturing. At the same time, the material and related processes are environmentally friendly and upon selecting the right electrolyte it yields into sustainable, environmentally friendly, supercapacitor design.

### Energy Density Boosted by Functionalized Group #1

The target of the activity was to assemble redox-active reversible materials (such as transition metal oxides -  $\text{MnO}_2$ , iron oxides or 2D MXenes) as a long functionalizing group perpendicular to the 2D graphene structure in order to create creating nano-pores and nano-channels boosting the charge storage mechanisms.

Metal iron tetraaminophthalocyanine G-FePc was selected as the best candidate matching the fluorographene manufacturing route using Bingel-Hirsch reaction<sup>2</sup>. (Figure 3.)

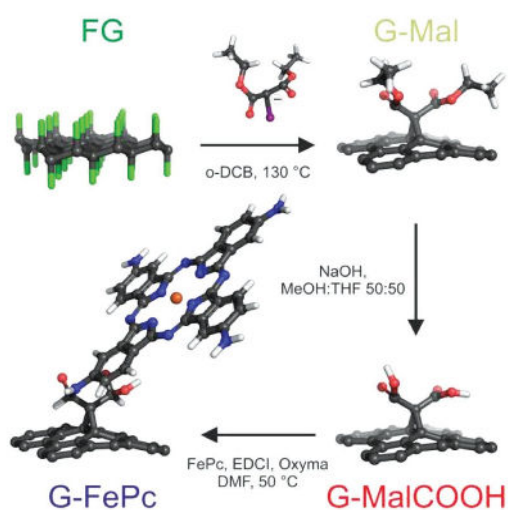


Figure 3. Covalent modification of graphene with FePc-NH<sub>2</sub> via the Bingel-Hirsch reaction on FG [2]

## Energy Density Boosted by Nitrogen Doping #2

The nitrogen doping of graphene via mild and low energy processes to afford homogeneous product composition and topology with high nitrogen content (>10 at. %) remains a challenge of contemporary 2D materials chemistry.

We reported a carbon-based electrode material, GN3, with an unprecedented density of  $2.8 \text{ g cm}^{-3}$  and an  $\text{N}_2$  sorption-based surface area of  $128 \text{ m}^2 \text{ g}^{-1}$  that can host ions even more efficiently than carbon materials with surface areas exceeding  $2000 \text{ m}^2 \text{ g}^{-1}$ . GN3, which is prepared by reacting graphite fluoride with sodium azide, has tetrahedral ( $\text{sp}^3$ ) C–C bonds, which were identified in the same region as the C–C bonds in diamond. However, it retains a 2-D structure with a very high content of aromatic (and thus conductive) regions, together with nitrogen superdoping in the vacancies and holes of the aromatic lattice.

This new class of carbon-based materials comprising nitrogen doped graphene with diamond-like tetrahedral bonds for high energy density supercapacitor electrodes. The cells with electrodes made from these materials offer ground-breaking energy storage capability at very high charging/discharging rates. The discovery of this class of materials will spur intense research on other high-density conductive carbon materials with different functionalities, with the aim of further increasing the competitiveness of supercapacitors in the portable energy storage landscape.

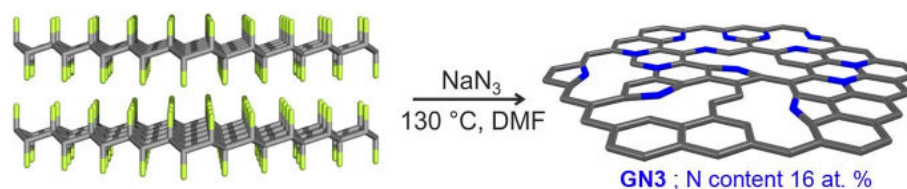


Figure 4. The schematic representation of a GN3 synthesis from fluorographene.

## Electrolyte Selection

Traditional electrolytes, some in production supercapacitors since 1999, are based on sulfuric acid or KOH **aqueous** solutions. They are RoHS and REACH, offer low internal resistance, but limit the voltage to roughly one volt per cell.

**Organic electrolytes** are generally based on acetonitrile or propylene carbonate that allow higher voltage per cell – a reason why they are currently favored but with higher internal resistance.

**Ionic electrolytes** are salts molten at room temperature. Compared to existing organic electrolytes, ionic liquids offer the potential for increased operating temperature, increased cycle life, improved safety, and increased energy density. By replacing conventional electrolytes with ionic liquids the energy storage capacity of ultracapacitors could double by allowing them to be charged to higher voltages. Ionic liquids proved to be the most promising electrolyte to enhance the voltage of supercapacitors to values higher than 3.0 V. However, their relatively high viscosity and low ionic conductivity reduce, at the present, the power performance of supercapacitors.

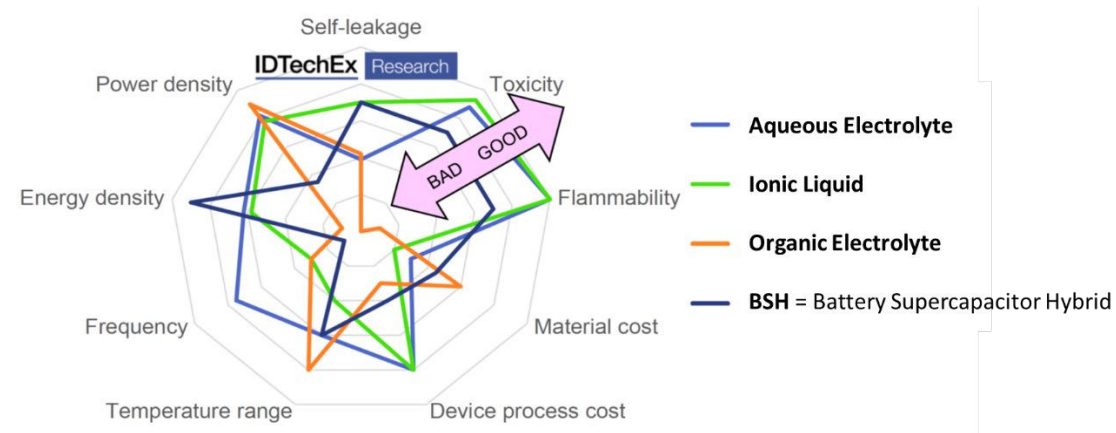


Figure 5. Electrolyte types benchmarking spider chart; used under permission from source: IDTechEx report<sup>3</sup>

## EXPERIMENTAL

Samples of functionalized graphene (Figure 6.) and N-doped graphene GN3 (Figure 7.) were prepared using fluorographene as the precursor. Details of chemistry and processes – see scientific publication “Nitrogen doped graphene with diamond-like bonds achieves unprecedented energy density at high power in a symmetric sustainable supercapacitor”<sup>4</sup>.

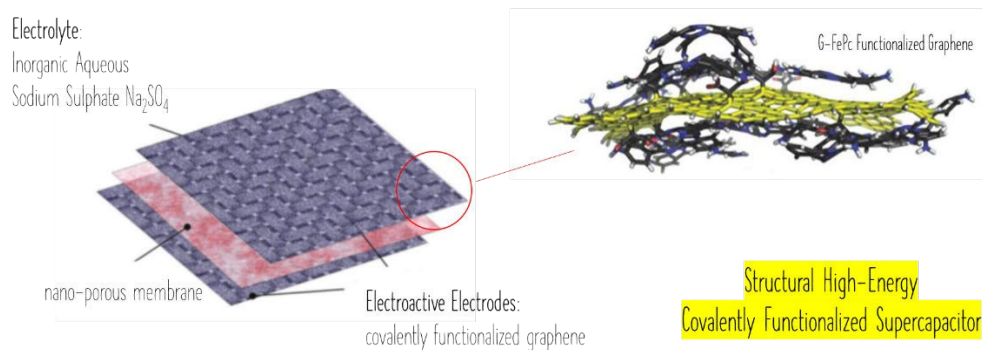


Figure 6. Functionalized graphene with G-FePc supercapacitor structure

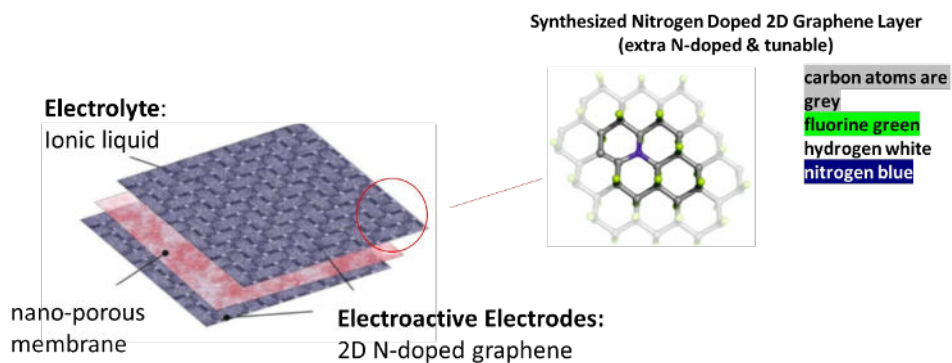


Figure 7. N-doped (GN3) 2D graphene supercapacitor structure

Electrodes prepared from active electrode powder were inserted into the capacitor contact holder and characterized using a measurement jig. Cycle lifetime at room temperature was also evaluated – see Figure 8.

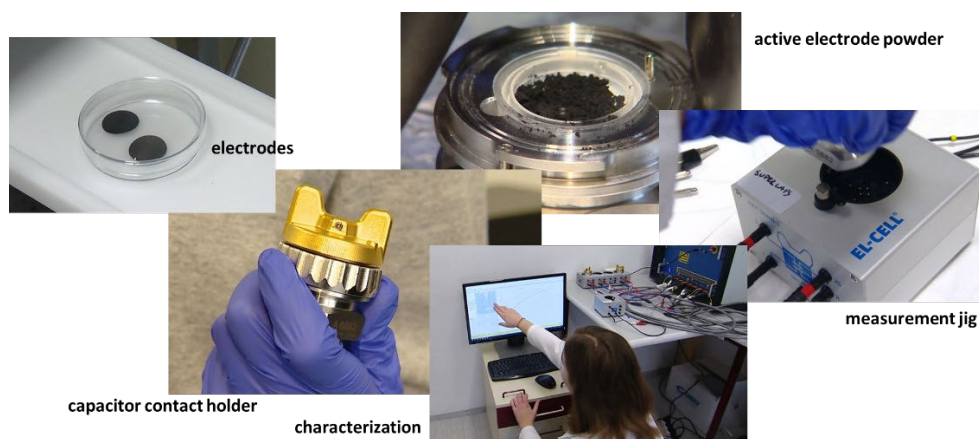


Figure 8. supercapacitor sample preparation and characterization

## RESULTS

Using this new approach, we achieved several important milestones for graphene-based supercapacitors as summarized below and in Figure 9.

### G-FePc Functionalized Graphene Supercapacitors:

- SC cell voltage 1V with an aqueous liquid
- Up to 1000 F g<sup>-1</sup> energy density at 59 Wh kg<sup>-1</sup> and power density of 0.3 kW kg<sup>-1</sup>
- Limiting factor: G-FePc high cost

### 2D N-Doped Graphene:

- SC cell voltage with ionic liquid: 3.7V (high cell voltage compared to conventional SCs)
- Up to 128 F g<sup>-1</sup> at ~55 Wh kg<sup>-1</sup> and 2 kW kg<sup>-1</sup> corresponds to 357 F cm<sup>-3</sup> at 170 Wh L<sup>-1</sup> and 5 kW L<sup>-1</sup> (lower gravimetric ED compared to covalently functionalized graphene). Importantly, this represents a record volumetric energy density, best in class ever reported for SCs (comparable to batteries).
- At higher power density (required for SCs), up to 300 F cm<sup>-3</sup> 140 Wh L<sup>-1</sup> and 52 kW L<sup>-1</sup> (record power density, significantly higher when compared to conventional SCs).
- 100% capacitance retention after 10 000 charge-discharge cycles

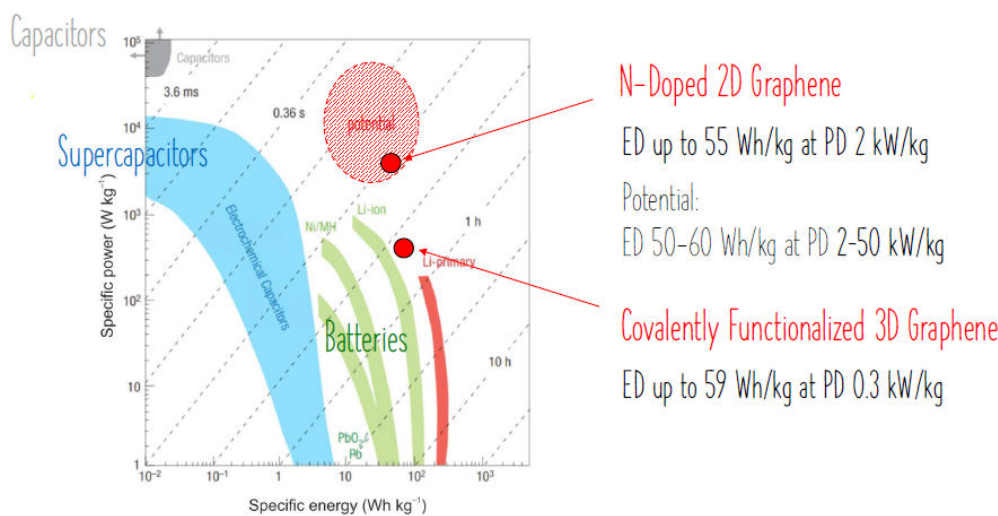


Figure 9. High energy SC achievements.

GN3 material, nitrogen super doped graphene with diamond-like layer bonds, achieved record energy density in a range of conventional batteries while power density is close to capacitors capability.

Figure 10. Shows state-of-the-art comparison of two GN3 tested materials: 1] GN3-lab, and 2] GN3-industrial (made by subcontractor following our recipe) with prominent analogues selected from the literature, recognized for their promising features; Holey graphene<sup>5</sup>, compressed microwave expanded and activated reduced graphene oxide<sup>6</sup>, tri-doped carbon<sup>7</sup>, capillary densified graphene<sup>8</sup>, ultrathick graphene<sup>9</sup>, MoS<sub>2</sub> in 1T phase<sup>10</sup>, and exfoliated graphene-mediated graphene oxide (EGM-GO)<sup>11</sup>.

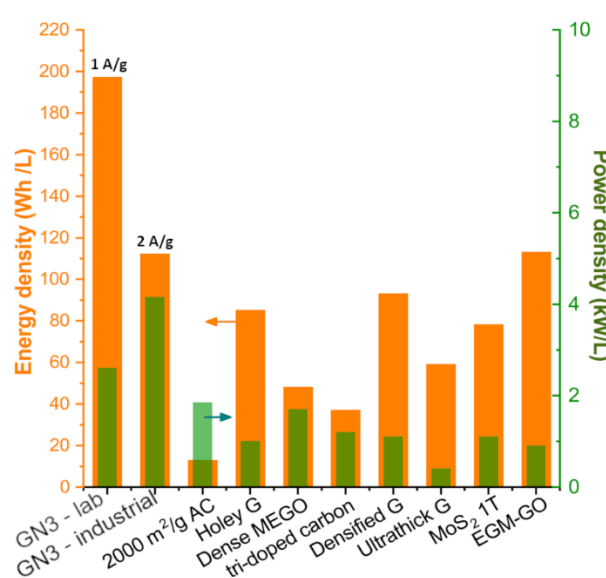


Figure 10. GN3 (our results) vs literature supercapacitors' ED and PD achievements.



## NEXT STEPS

### 1. Mass Production Ready High Energy Graphene-Based Supercapacitors

The next step will be to build prototypes of supercapacitors in co-operation with international partners. We will focus on optimizing the properties of GN3 material and move to pilot production of the graphene-based electrode supercapacitors.

The aim is to increase the energy density of supercapacitors beyond  $50 \text{ Wh L}^{-1}$ , which will allow their wide use in electric vehicles and as battery supports in devices that need to be supplied with large amounts of energy in a very short time.

For the commercial application, the next step includes the preparation of SC prototypes by our industrial partners in a flat pack and wound packaging form to demonstrate its capabilities in the most common SC forms on the market today.

### 2. Use of Fluorographene Process to Enhance LiS Batteries

The fluorographene process has also the potential to enhance the electrode system in LiS batteries.

Sulfur represents a low-cost, lighter weight, sustainable, and high theoretical capacity cathode material for lithium–sulfur batteries. However, the shuttling-effect of the formed lithium polysulfides, as well as their low conductivity, compromise the electrochemical performance of lithium–sulfur cells.

Li-S batteries advantage:

- 3-5x higher energy density  $\sim 2,500 \text{ Wh/kg}$
- Nickel and cobalt free solution (conflict/resource issues free)
- No heavy metals / 3x lighter compared to Li-Ion
- No / suppressed thermal runaway failure
- Cheaper:  $\sim 200$  tons of sulfur equals to  $\sim 1$  ton of cobalt
- Lower carbon footprint, RoHS friendly
- 3x Faster charging
- No extra pressure needed (unlike solid states)
- Vigorous 100% charge/discharge depth rate

To tackle this challenge, a so far unexplored cathode, composed from sulfur covalently bonded directly on graphene is developed. This is achieved by leveraging the nucleophilicity of polysulfide chains, which react readily with the electrophilic centers in fluorographene, as experimental and theoretical data unveil. The reaction leads to the formation of carbon–sulfur covalent bonds and a particularly high sulfur content of 80 mass%.

The first results of Li-S batteries with enhanced electrodes by fluorographene process keep the capacity at  $\approx 700 \text{ mAh g}^{-1}$  after 100 cycles at 0.1 C, at  $\approx 644 \text{ mAh g}^{-1}$  after 250 cycles at 0.2 C and  $\approx 470 \text{ mAh g}^{-1}$  even after 500 cycles that present promising results for further optimization and development.

## SUMMARY AND CONCLUSION

The presented research led to the development of new high energy graphene-based electrode material and preparation processes for a new generation of supercapacitors (SC) as highly efficient, high-energy density and sustainable storage components.

The nitrogen-doped GN3 graphene material that has been developed is proving to be very promising for use in supercapacitors. The material has a higher density than graphite, which, combined with the great ability to host ions from the electrolyte, leads to very high volumetric energy density, i.e., the amount of energy stored per unit of volume of the material. This is significantly higher than for all carbon- or graphene-based supercapacitor materials described so far, which can bring a breakthrough improvement in the performance of supercapacitors.

The GN3 based structure supercapacitors demonstrated herein have reached energy density as high as  **$200 \text{ Wh L}^{-1}$  and power density up to  $50 \text{ kW L}^{-1}$**  that presenting a record achievement reported at the time of the research. Laboratory made supercapacitors confirmed excellent stability of the material, no/low wear out during charge / discharge cycle and

long lifetime of the system. Further optimization of the GN3 material for prototyping and pre-production is under process to prepare for mass production of the novel graphene SCs in an industrial scale.

#### **Achievement summary:**

- New generation of graphene-based supercapacitor electrode material with high mass density
- Record-high volumetric energy and power densities
- Patent application filed
- Scalable synthesis ready for pre-production

#### **ACKNOWLEDGEMENT**

##### **The research has been supported by:**

European Union's Horizon 2020: ERC project No. 683024

European Union's ERC Grant: ERC PoC 899245

European Regional Development Fund: Operational Programme Research, Development and Education CZ.02.1.01/0.0/0.0/16\_019/0000754

Ministry of Education, Youth and Sports of the Czech Republic: LO1305, CZ.1.05/2.1.00/19.0377, Research Infrastructure NanoEnviCz: project No. LM2015073

Internal Student Grant Agency of the Palacký University in Olomouc, Czech Republic (IGA\_PrF\_2021\_031, IGA\_PrF\_2022\_019).

#### **IP PROTECTION**

EP 20173178.3 – Title: Nitrogen doped graphene and use thereof

PCT/CZ2021/050016 – Follow-up Application

EP 20199904.2 – Title: Sulfur-functionalized graphene, and use thereof as Li-S battery cathode

PCT/CZ2021/050045 – Follow-up Application

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